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Chapter 20: Acids and  
Bases. STUDY.

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Terms in this set (27)

acid. a compound that  
produces hydrogen

ions in a solution. acid

dissociation constant.

the ratio of the

concentration of the

dissociated form of an

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acid to the undissociated form.

## **Chapter 20: Acids and Bases**

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- Acids turn blue litmus paper red, taste sour, and react with metals to produce hydrogen gases.
- Common acids include vinegar ( $\text{HC}_2\text{H}_3\text{O}_2$ ) and stomach acid ( $\text{HCl}$ ).
- Bases turn red litmus paper blue, have a bitter taste, and are slippery to the

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touch. • Common bases include Drano ( $\text{NH}_4\text{OH}$ ).

## **Chapter 20 Acids and Bases - studylib.net**

Chapter 20 acids and bases. STUDY. PLAY. Hydrogen-ion donor. an acid, according to the Bronsted-Lowry theory. Hydrogen-ion acceptor. A base, according to the Bronsted-Lowry theory, ammonia acts as a base when it

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accepts hydrogen ions from water. Conjugate acid.

## **Chapter 20 acids and bases**

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Chemistry Chapter 20  
Acids and Bases. Acids.  
Bases. Acids. Base.

Taste sour, turns litmus  
paper red, reacts with  
some metals to....

Tastes bitter, feels  
slippery, turns litmus  
paper blue, reacts....

compounds that form

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hydrogen ions when dissolved in water. A compound that produces hydroxide ions when dissolved in water.

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Bases, and pH  
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ACIDS & BASES All  
acids and bases are  
solutions. Therefore, a  
solid understanding of  
Solution Chemistry and  
Equilibrium will  
guarantee you success  
in this unit. I. Defining

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Acids & Bases: A.  
General Characteristics  
of Acids & Bases:  
Properties of Acids  
Properties of Bases  
Ionic compounds which  
\_\_\_\_\_ in H<sub>2</sub>O to give a  
solution that: Ionic  
compounds which  
dissociate in H<sub>2</sub>O to  
give a ...

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The Brønsted-Lowry definitions of acids and bases are that acids are proton (hydrogen ion) donors and bases are proton acceptors. All Arrhenius acids and bases are included in

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the definitions of Brønsted-Lowry definitions, but there are some Bronsted-Lowry acids or bases that are not Arrhenius acids or bases. For example ammonia ( $\text{NH}_3$ )

## **CHEMISTRY NOTES - CHAPTERS 20 AND 21 Acids and Bases**

...

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University of California,  
San Diego. The  
Properties of Acids and  
Bases Chapter 20 Acids  
and bases Arrhenius  
definition of

## **Chapter 20 The Properties of Acids and Bases BB3 - The**

...

Figure 20.1 All of these  
items contain acids or  
bases, or produce acids  
or bases upon  
dissolving in water.

Tomatoes contain

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ascorbic acid; tea contains tannic acid. The base calcium hydroxide is a component of mortar, which was used to build the Great Wall of China. Antacids use a variety of bases to neutralize excess stomach acid.

## **Chapter 20 | ACIDS AND BASES - physci.us**

Chapter 20: Acids,  
Bases, and Redox. 20.1

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Acid and bases; 20.2  
Relative Strengths of  
Acids and Bases; 20.3  
Acidic, Basic, and  
Neutral Solutions; 20.A  
Acid Rain and Basic  
Oceans; 20.4 Losing  
and Gaining Electrons;  
20.5 Electrochemistry;  
20.6 Corrosion and  
Combustion

## **Chapter 20: Acids, Bases, and Redox | Conceptual Academy**

In chemistry, acids and  
bases have been



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defined differently by three sets of theories: One is the Arrhenius definition defined above, which revolves around the idea that acids are substances that ionize (break off) in an aqueous solution to produce hydrogen ( $H^+$ ) ions while bases produce hydroxide ( $OH^-$ ) ions in solution.

## **16: Acid-Base Equilibria - Chemistry**

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$\text{HX}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{X}^-(\text{aq})$

Acid Base Conjugate

Conjugate acid base

Conjugate acids and

bases The forward

reaction is the reaction

of an acid and a base.

of an acid and a base.

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The reverse reaction is also the reaction of an acid and a base.

### **Chapter 18: Acids and Bases**

Title: Chapter 10: Acids and Bases 1 Chapter 10 Acids and Bases.

When we mix aqueous solutions of ionic salts, we are not mixing single components, but rather a mixture of the ions in the solid ; The ionic solid dissolves in the water ; We call a

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compound that  
dissolves in water  
soluble and if it doesn't,  
it is insoluble; 2  
Electrolytes

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